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Atoms of a given element are identical in size, mass and other properties; atoms of different elements differ in size, mass, and other properties. 3. Atoms cannot be subdivided, created or destroyed. 4. Atoms of different elements combine in simple whole-number ratios to form chemical compounds. 5.

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in molecules containing the

3 Atoms: The Building Blocks of Matter

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in molecules containing the

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CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

5 The Periodic Law

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State the following measured quantities in the units indicated. a. 5.2 cm of magnesium ribbon in millimeters b. 0.049 kg of sulfur in grams c. 1.60 mL of ethanol in microliters d. 0.0025 g of vitamin A in micrograms e. 0.020 kg of tin in milligrams f. 3 kL of saline solution in liters

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