

## Math Skills Transparency Worksheet Answers Chapter 14

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Answers

MATHSKILLSTRANSPARENCY WORKSHEET Using the Periodic Table 1. Identify the number of valence electrons in each of the following elements. a. Ne d. Mg f. Cl h. Si 2. Identify the energy level of the valence electrons in each of the following

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elements. a. c. d. e. H Ar 1 3.

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MATH SKILLS TRANSPARENCY WORKSHEET Determining Numbers of Ions Use with Chapter 7, Section 7.3 1. Write a simple word equation that illustrates what must be true of total positive charge and total negative charge in an ionic compound. 2. How many potassium ions (group 1) would be related to balance the charge of each of the following in a compound?

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Math Skills Transparency Worksheets Chemistry: Matter and Change Chapter 7 ...

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T102 Chemistry: Matter and Change Name Date Class Name Date Class MATH SKILLS TRANSPARENCY WORKSHEET 1 MATH SKILLS TRANSPARENCY WORKSHEET 2 Interpreting and Drawing Graphs Use with Chapter 2, Visualizing the Use with Chapter 3, Section 2.4 Conservation of Mass Section 3.2 1.

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Math Skills Transparency 2 Worksheets - Kiddy Math

The exponents are added.  $(2.00 \times 10^4) \times (1.00 \times 10^3) = 2.00 \times 10^{4+3} = 2.00 \times 10^7$  Division When numbers in scientific notation are divided, only the number is divided. The exponents are subtracted.  $9.60 \times 10^7 \div 1.60 \times 10^4 = 6.00 \times 10^{7-4} = 6.00 \times 10^3$  Chemistry: Matter and Change Name Date Class 58 MAT" HANDBOOK TRANSPARENCY WORKSHEET Operations with Scientific Notation I. Perform the following operations and express the answers in scientific notation.

Mr. Bastien - Home

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MATH SKILLS TRANSPARENCY WORKSHEET Use with Chapter 1, Section 11.1 MATH SKILLS TRANSPARENCY WORKSHEET 16  
Section 11.2 Determine the mole ratios for each of the following balanced chemical equations.  $1.2\text{C}(s) + \text{O}_2(g) \rightarrow 2\text{CO}(g)$  1  
mol  $\text{O}_2$  2 mol C 2 mol  $\text{CO}$  / 2 mol C 2 mol C / 1 mol  $\text{O}_2$  2 2 mol  $\text{CO}$  / 1 mol  $\text{O}_2$  2 2 mol C / 2 mol  $\text{CO}$  1 mol  $\text{O}_2$  / 2 mol  $\text{CO}$  2 2.  $\text{WO}_3$   
(s) +  $3\text{H}_2$  (g)  $\rightarrow \text{W}(s) + 3\text{H}_2\text{O}(l)$

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