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First Law of Thermodynamics

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□ This law states that (1) heat is a form of energy that (2) its conversion into other forms of energy is such that total energy is conserved. □ The change in the internal energy of a system is equal to the heat



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added to the system minus  
the work done by the system:

**Ideal Gas Law (Equation of  
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heat, work, and temperature,  
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be able to state the First Law and to define heat, work, thermal efficiency and the difference between various forms of energy.

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established principle of thermodynamics (which eventually became the Second Law) was formulated by Sadi Carnot in 1824. By 1860, as found in the works of those such as Rudolf Clausius and William Thomson, there were



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