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whether the sentence or
statement is true or false.
1. The region in space where
an electron is most likely
to be found is called an
energy level. 2. Electron

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configurations are often condensed by writing them using the previous noble gas core as a starting point.

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~~Welcome to Dr. Mintz's~~
~~Chemistry Classroom Home~~
Solution: 0.50 mmol
 $\text{mL} \times 25.0 \text{ mL} = 12.5 \text{ mmol HCl(aq)}$
 $\text{HCl(aq)} \text{ [HCl(aq)] } 12.5 \text{ mmol}$
 $= 0.10 \text{ mmol HCl(aq)}$
 $/\text{mL} \times 120 \text{ mL} = 12.0 \text{ mmol}$ 2. (a) Given:

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$V \text{ HCl(aq)} = 25.00 \text{ mL};$

$[\text{HCl(aq)}] = 0.350 \text{ mol/L};$ V

$\text{NaOH(aq)} = 5.0 \text{ mL};$

$[\text{NaOH(aq)}] = 0.500 \text{ mol/L};$

Required: $[\text{H}^+(\text{aq})]$, pH

Analysis: Before titration,

amount of H^+ is: $\text{H}^+(\text{aq})$ $\text{H}^+(\text{aq})$

$\text{H}^+(\text{aq})$ $0.350 \text{ mmol mL} \times 25 =$

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nc V

~~Section 8.7: Acid-Base
Titration Tutorial 1
Practice, page 547~~

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Identify any errors in the

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structure by drawing them.
Rename the structure
correctly. 2-ethyl-2,4,4-tri
methylpentane What Is
Required? You are to
identify the errors in the
name by drawing the
structure and renaming it

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correctly. What Is Given?
You are given an incorrectly
named alkane.

~~Unit 1 Organic Chemistry~~
~~Mr. Arthur's Science Page~~
= 0.12 mol/L Statement: The
equilibrium concentration of

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dinitrogen tetroxide gas is
0.12 mol/L. 2. Given:
initial quantity of NOCl(g)
= 3.00 mol; volume = 3.00 L;
 $[\text{NO(g)}]$ initial = 0.00
mol/L; $[\text{Cl}_2\text{(g)}]$ initial =
0.00 mol/L; $[\text{NO(g)}]$
equilibrium = 0.043 mol/L

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Required: $[\text{NOCl}(\text{g})]$
equilibrium; $[\text{Cl}_2(\text{g})]$
equilibrium Solution: Step
1.

~~Section 7.1: Equilibrium
Systems Mini Investigation:
The ...~~

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Solution: bonds broken 3 mol
1 mol CH CCl 3 mol HnD DD --
 $\Delta = x = x + x = \sum 413 \text{ kJ mol}^{-1} \times 3 + 1$
 $\text{mol} \times 339 \text{ kJ mol}^{-1} \times \Delta H = 1578 \text{ kJ}$
Statement: The energy
required to separate 1 mol
of chloromethane into free
atoms is 1578 kJ. 2.

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Solution: Step 1: Write the balanced chemical equation for the combustion of 1 mol of ethene gas. $C_2H_4(g) + 3O_2(g) \dots$

~~Section 5.3: Bond Energies
Tutorial 1 Practice, page~~

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Solution: $q=Ne$

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$= (1.2 \times 10^{-8}) (1.602 \times 10^{-19} \text{ C})$
 $q = 1.92 \times 10^{-11} \text{ C}$ (one extra
digit carried) $F = k \frac{q_1 q_2}{r^2} = 8.99 \times 10^9 \text{ N} \cdot \text{m}^2 / \text{C}^2$
 $F = \frac{(1.92 \times 10^{-11} \text{ C})^2}{(1.10 \text{ m})^2}$
 $F = 2.7 \times 10^{-12} \text{ N}$ Statement:
The force of repulsion

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between the two plastic spheres is 2.7×10^{-12} N. 2.
Given: $m = 2.48 \times 10^{-15}$ kg;
 $\Delta d = 1.7$ cm = 0.017 m; ΔV b
= 260 V; $e = 1.602 \times 10^{-19}$
C; $g = 9.8$ m/s² Required: q ;
N

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