

Chapter 8 Covalent Bonding Study Guide Answers

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Chem chapter 8 lecture – covalent bonding HW: read chapter 8.1 and make a vocab list The covalent Bond-Why do atoms bond Atoms bond to gain stability within their electrons-In covalent substances, atoms share electrons to achieve this-Atoms are still trying to follow the octet rule o Apart from hydrogen which follows the duet rule (wants 2 valence electrons) What is a covalent bond?

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Draw the Lewis structure for each molecule. 1. PH_3 H_2O H_2S H_2O_2 H_2O_3 respectively, for single, double, and triple P — — 2. H_2S H_2O H_2S_2 ...

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Chapter 8: Covalent Bonding 8.1 Covalent Bonding • covalent bond is a shared electron pair that connects the atoms in molecular (covalent) compounds • when two atoms get close enough, their electron clouds overlap, allowing the nucleus to attract the other atom's electrons which results in a net attraction between the two atoms • if atoms are too close together, they will repulse each other 8.2 Single Covalent Bonds and Lewis Structures • shared electron pairs occupy same shell as ...

8 - Covalent Bonding - Chapter 8 Covalent Bonding 8.1 ...

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242 Chapter 8 • Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

Chapter 8: Covalent Bonding - Madison County School District

Chemistry: The Central Science Chapter 8: Basic Concepts of Chemical Bonding The properties of substances are determined in large part by the chemical bonds that hold their atoms together 8.1: Chemical bonds, Lewis Symbols, and the Octet Rule Chemical bond – The attraction that causes two atoms or ions to be strongly attached to each other o Ionic bond – electrostatic forces that exist between ions of opposite charge o Covalent bond – bond formed by the sharing of electrons between two ...

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A covalent bond in which one atom contributes both bonding electrons Polyatomic ion A tightly bound group of atoms that has a positive or negative charge and behaves as a unit.

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