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2. Gas particulars are in random, constant, straight line motion. 3. When particles collide with each other (or any container), the collisions are said to be elastic. 4. The volume that gas particles take up is negligible. And from the distance between particles is relatively great.

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BEHAVIOR OF GASES REVIEW Page 102 Chemistry Unit Assessment 2007 Baltimore County Public Schools 11. Calculate the new temperature of a gas when 1500 mL at 25oC is suddenly compressed to 500 mL. Charles ' Law K mL K mL V T V T 100 (1500) (298)(500) 1 1 2 2 12. A flask contains 34.6 kPa of CO 2

Student Review Packet Answer Key

Behavior Of Gases Review 2 SECTION 2 BEHAVIOR OF GASES 1. a measure of how fast the particles of an object are moving 2. when it is heated 3. Temperature of gas particles Energy of gas particles Volume of gas particles 1) 20 ° C Particles have the smallest amount of energy. Volume is smallest. 2) 50 ° C Particles have more

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Behavior Of Gases Review 2 Answers properties of gases. The imprecision is known as the non-ideal behavior of gas, and the van der Waals equation. (P + n2a V2)/(V - nb) = nRT. has been introduced to deal with non-ideal behavior of gases in Ideal gas law. Gases - A Review - Chemistry LibreTexts Behavior Of Gases Review 2 Answers Page 8/26

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You usually cannot feel it, but air has pressure. The gases in Earth ' s atmosphere exert pressure against everything they contact. The atmosphere rises high above Earth ' s surface. It contains a huge number of individual gas particles. As a result, the pressure of the tower of air above a given spot on Earth ' s surface is substantial.

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